SYNTHESIS AND ANTHELMINTIC ACTIVITY OF FLUORO BENZOTHIAZOLE LINKED THIAZOLIDINONE

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ABSTRACT

Fluorobenzothiazole linked Thiazolidinones using 3-nitro benzaldehyde, various novel derivatives were synthesized and evaluated for anti-microbial, anti-tubercular, anthelmintic activity. Structures of these compounds has been established by melting point, elemental analysis, UV and IR data. Compounds showed significant anthelmintic activity. **KEY WORDS:** Anthelmintic activity, Fluorobenzothiazole, Thiazolidinones.

1.INTRODUCTION

Benzothiazoles incorporated with fluorine (Feller, 1995) get promising molecule for clinical to use.Fluorobenzothiazole (Trivedi and Shah, 1992) are biologically versatile as they have been found to be anticancer agent anti-tubercular agent (Litvinchuk and Farmakol, 1963), fungicidal (Rao, 1996), cardiovascular agent, local anaesthetic activity. The above said nucleus is modified by linking thiazolidinones, the derivatives of thiazolidine which belong to the important group of heterocyclic compounds (Satizinger and Liebijs, 1963). In recent years several new methods for the preparation of thiazolidinone derivatives and reactions have been reported in literature. This attempt is made to link fluorobenzothiazole with potent thiazolidinone in present work by using 3nitrobenzaldehyde in hope to get promising clinically useful biodynamic compounds.

2.MATERIALS AND METHODS

Purity of compounds was checked by TLC. Melting points were determined by open capillaries method and are uncorrect. IR spectra (NaCl) are recorded on FTIR (Schimadzu-84005) spectrophotometer using nujol mull technique. For anthelmintic activity *in vitro* technique by using earthworms species *Perituma posthuma*, (Jayachandran, 2003). Compounds showed significant anthelmintic activity.

EXPERIMENTAL

First Step Synthesis of 2-amino-6-fluoro-7chloro(1,3)benzothiazole (1)

To the glacial acetic acid (20ml) which is cooled below room temperature,8gm (0.08mol) of potassium

thiocyanate and 1.45g (0.01 mol) of fluorochloroaniline was added. The mixture was placed in freezing mixture of ice and salt, mechanically stirred while 1.6ml of bromine in 6ml of glacial acetic acid was added, from a dropping funnel at such a rate that the temperature never raised beyond room temperature. After all the bromine was added (105min), the solution was stirred for 2 hours below room temperature and at room temperature for 10 hours, it was then allowed to stand over night, during which period an orange precipitate settle at the bottom, water (6ml) was added quickly and slurry was heated at 85°c on a steam bath and filtered hot. The orange residue was placed in a reaction flask and treated with 10ml of glacial acetic acid, heated again to 85°c and filtered hot. The combined filtrate was cooled and neutralized with concentrated ammonia solution to $p^H 6$. A dark yellow precipitate was collected. Recrystalised from benzene, ethanol (1:1), after treatment with animal charcoal gave yellow plates of 2-amino-6-fluoro-7chloro-(1,3) benzothiazole. After drying in a oven at 80°c, the dry material (1gm, 51.02%) melted at 210212°c. UV 307.4, 269nm, IR 1542cm⁻¹(aromatic C=C) and 3475cm⁻¹ (NH₂), 1456 cm⁻¹(thiazole), 1215 cm⁻¹(aromatic-F), 712 cm⁻¹(aromatic-Cl).

Second Step Synthesis of 2-[m-nitrobenzylidine]-6fluoro-7chloro (1, 3) benzothiazole (2)

0.01 mol of 2-amino-6-fluoro-7-chloro (1, 3) benzothiazole with 0.015 mol solution of m-nitro benzaldehyde, added 20 ml ethanol and 3-4 drops of HCl and refluxed for 2-3 Hrs. Solution cooled and poured into crushed ice. Recrystalised with benzene and ethanol.

Third Step Synthesis of 3-[6'-fluoro-7'-chloro(1',3') benzo thiazol-2'-yl]m-nitrophenyl(1,3) thiazolidine-4-one (3)

A mixture of Schiff's base (0.01 mol) and 0.025 mol of 2-thioglycolic acid heated on oil-bath at $115^{\circ}-120^{\circ}$ c for 12 Hrs. After reflux cool and triturated with 10% sodium bicarbonate solution. Crystallized from water. **Preparation of various derivatives (BT1-BT9)** 3-[6'-fluoro-7'-chloro(1',3') benzo thiazol-2'yl]mnitrophenyl(1,3) thiazolidine-4-one were treated with equimolar quantities of various aromatic amines, refluxed for 2 hours in presence of DMF, recrystalised from alcohol and benzene.

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SCHEME



Table No. 1 Analytical Data of the Compounds (BT₁-BT₉)

Compds	R	м.р (°С)	Yield (%)	Molecular Formula	Molecular Wt.	Elemental Analysis Data (Calculated in %)			
						С	н	N	
BT1		115	65.00	$C_{22}H_{14}N_5O_5S_2F$	511	51.66	2.73	13.69	
BT_2	-	119	76.00	$C_{22}H_{14}N_5O_5S_2F$	511	51.66	2.73	13.69	
BT3	-C>-****	101	72.00	$C_{22}H_{14}N_5O_5S_2F$	511	51.66	2.73	13.69	
BT4	Ő	182	69.00	$C_{22}H_{14}N_4O_3S_2FC1$	501	52.69	2.79	11.17	
BT ₅	-	202	52.00	$\mathbf{C}_{22}\mathbf{H}_{14}\mathbf{N}_4\mathbf{O}_3\mathbf{S}_2\mathbf{F}\mathbf{C}\mathbf{I}$	501	52.69	2.79	11.17	
\mathbf{BT}_{6}		152	63.00	$C_{22}H_{14}N_4O_3S_2FCI$	501	52.69	2.79	11.17	
\mathbf{BT}_7	— NO	174	58.00	$C_{20}H_{17}N_4O_4S_2F$	406	52.17	3.69	12.17	
BT8		187	56.00	$C_{20}H_{18}N_5O_3S_2F$	459	52.28	3.92	15.25	
BT ₉	N (C ₀ H ₀) ₂	118	63.00	$C_{28}H_{19}N_4O_3S_2F$	542	61.25	3.50	10.33	

Table2. IR spectral assignments of synthesized compounds (BT₁-^{BT}₉)

	Characteristic absorption bonds (in cm ⁻¹)											
Compounds	Ar- NH ₂ Str.	C=0 Str.	Aro.C=C Str.	C-F Str.	C- Cl Str.	NO ₂	C=C	3º- Nitrogen	С-Н	C-S- C		
\mathbf{BT}_1	3385	1770	1597	1260		1370	1597	3020	3460	1301		
BT ₂	3400	1870	1625	1249		1309	1625	3090	3433	1392		
BT ₃	3390	1820	1697	1249		1303	1525	3050	3300	1392		
BT4	3128	1790	1595	1195	1170		1670	3080	3380	1352		
BT5	3228	1820	1670	1219	1166		1525	3093	3352	1307		
BT ₆	3201	1825	1690	1249	1197		1525	3095	3435	1360		
BT ₇	3200	1800	1550	1295			1380	3360	3240	1336		
BT ₈	3370	1825	1600	1250			1590	3100	3460	1300		
BT ₉	3350	1835	1597	1290			1395	3020	3400	1330		

Procedure for Anthelmintic activity (Jayachandran, 2003)

synthesized compounds are screened for The anthelmintic activity by using earthworms. Six earthworms of nearly equal size were placed in standard drug solution and test compounds solutions at room temperature. Normal saline was used as control. The standard drug and test compounds were dissolved in minimum quantity of DMF and the volume was adjusted upto 15 ml. with normal saline solution to get the concentration of 0.1%, 0.2% and 0.5%. Albendazole was used as a standard drug. The compounds were evaluated by the time taken for complete paralysis and death of earthworms. The mean lethal time for each test compound was recorded and compared with standard drug. The time taken by worms to become motionless was noted as paralysis time. To ascertain the death of the motionless worms, these were frequently applied the external stimuli, which stimulates and induces movement in the worms, if alive.

The mean lethal time and paralysis time of earthworms for different test compounds and standard drug are tabulated in Table 3.

Table No. 3 Anthelmintic activity (BT₁-BT₉)

	Concentration (%)								
Compounds	Time fo	or paralysis	Time for death (min.)						
	0.1	0.2	0.5	0.1	0.2	0.5			
Control									
Albendazole	50	45	40	70	63	55			
BT ₁	48	33	30	130	125	110			
BT ₂	39	45	47	97	102	98			
BT ₃	23	25	20	130	129	118			
BT ₄	26	20	19	110	94	85			
BT ₅	48	35	30	192	168	128			
BT ₆	49	38	38	174	168	120			
BT ₇	32	30	26	125	118	110			
BT ₈	45	34	28	130	125	101			
BT ₉	60	42	32	179	157	122			

3.RESULTS AND DISCUSSION

Some novel analogs of the fluorobenzothiazole schiffs base using m-nitro benzaldehyde and screened for anthelmintic activity using *Perituma posthuma* earthworms. Some of the compounds showed significant activity.

REFERENCES

Feller R J, Fluorine Chem., 33, 1995, 366.

Jayachandran E, Nargund L.V.G, Roy A, Indian Drugs, 40(7), 2003.

Litvinchuk M.D, Farmakol, Chem.Abstr., 60, 1963, 13761e.

Rao R.P, Curr.Sci., 35, 1996, 54.

Satizinger G, Liebijs, Ann.Chem., 665, 1963, 150.

Trivedi and Shah B, Chem.Abstr., 116, 1992, 151637c.